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QUESTIONS © 2014 The College Board. Visit the College Board on the Web: [www.collegeboard.org](http://www.collegeboard.org). GO ON TO THE NEXT PAGE. -7-  $CH_3CH_2COOH(aq) + H_2O(l) \rightleftharpoons CH_3CH_2COO^-(aq) + H_3O^+(aq)$  2. Propanoic acid,  $CH_3CH_2COOH$ , is a carboxylic acid that reacts with water according to the equation above. At

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(d) Calculate the number of moles of precipitate that is produced in the experiment.  $1.698 \text{ g} - 1.462 \text{ g} = 0.236 \text{ g}$   
 $\text{PbI}_2 (\text{s}) \quad 2 \quad 4 \quad 22 \quad 2 \quad 1 \text{ mol PbI}_2 \quad 0.236 \text{ g PbI}_2 = 5.12 \cdot 10^{-4} \text{ mol PbI}_2$

[AP Chemistry 2014 Scoring Guidelines - College Board](#)  
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in solution at the beginning of the reaction and at the end, do not change in any way (i.e., they are spectators), and as such can be omitted from the equation.

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